

Chapter 10 Supplemental Problems Chemical Reactions Answers

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Chapter 10 Supplemental Problems Balance the following chemical equations. 1) $\text{SnS}_2 (\text{s}) + \text{O}_2 (\text{g}) \rightarrow \text{SnO}_2 (\text{s}) + \text{SO}_2 (\text{g})$ 2) $\text{C}_2\text{H}_6 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{g})$ 3) $\text{Al} (\text{s}) + \text{HCl} (\text{aq}) \rightarrow \text{AlCl}_3 (\text{aq}) + \text{H}_2 (\text{g})$ 4) $\text{CoCO}_3 (\text{s}) \rightarrow \text{CoO} (\text{s}) + \text{CO}_2 (\text{g})$ Write a balanced equation for each of the following reactions, substituting symbols and formulas for names.

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Chapter 10 Supplemental Problems The Supplemental Problems for Chapter 10. Supplemental Problems for Chapter 10. 1) One overarching trend in atomic structures is that as the coordination number increases, the interatomic separation increases. Using the appropriate equations, answer the following questions. Supplemental Problems for Chapter 10

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